

## 17.5

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                              |                    |
|-------------------|------------------------------|--------------------|
| 1. autoionization | 2. conjugate acid of a base  | 3. pH              |
| 4. autoionization | 5. conjugate base of an acid | 6. polyprotic acid |
|                   | 7. hydronium ion             |                    |

## Key Equations

- |                                    |                                                                                  |
|------------------------------------|----------------------------------------------------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the \_\_\_\_\_ of the conjugate base of a weak acid?
2. \_\_\_\_\_ is the conjugate base of the conjugate acid of a weak acid.
3. \_\_\_\_\_ is the conjugate acid of a weak acid.
4. \_\_\_\_\_ is the conjugate base of a weak acid.
5. \_\_\_\_\_ is the conjugate acid of a weak acid.
6. \_\_\_\_\_ is the conjugate base of a weak acid.
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17. \_\_\_\_\_ is the conjugate acid of a weak acid.
18. \_\_\_\_\_ is the conjugate base of a weak acid.
19. \_\_\_\_\_ is the conjugate acid of a weak acid.
20. \_\_\_\_\_ is the conjugate base of a weak acid.

1. \_\_\_\_\_
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# 19 2 Hydrogen Ions And Acidity Section Review

**Ian Pickup**



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