

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations, pH, and pOH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acids.

## Vocabulary

- |                   |                                       |        |
|-------------------|---------------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair           | 3. pH  |
| 4. hydronium ion  | 5. acid ionization constant ( $K_a$ ) | 6. pOH |
| 7. hydroxide ion  | 8. weak acid                          |        |

## Key Equations

- |  |  |
|--|--|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pOH} = -\log[\text{OH}^-]$                                 | 4. $\text{pH} + \text{pOH} = 14$             |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Write the number in the space provided.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
8.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)
10.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (acid-base reaction)

1. conjugate acid
2. conjugate base
3. conjugate acid
4. conjugate base
5. conjugate acid
6. conjugate base
7. conjugate acid
8. conjugate base
9. conjugate acid
10. conjugate base

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