

### 18.4 Acid-Base Chemistry: Theory

#### Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance  
 (B) volume, molarity, pH, amount of substance  
 (C) volume, molarity, pH, concentration of ions  
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base reaction is represented by the equation  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ . The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ . The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ .

(A)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (B)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

(C)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (D)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (E)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

3. The reaction  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$  is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ . The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ .

(A)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (B)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

4. Which of the following is a correct statement? (correct)

(A)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (B)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (D)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$       (E)  $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$       (F)  $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

5. A solution of  $\text{HCl}$  and  $\text{NaOH}$  is prepared. The pH of the solution is 1. The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ . The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ .

(A)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (B)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

(C)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (D)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (E)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

6. The reaction  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$  is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ . The reaction is exothermic. The enthalpy change for the reaction is  $-57 \text{ kJ mol}^{-1}$ .

(A)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$       (B)  $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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**David Baud**



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