

Electrode	Reaction	SRP (at 298 K)
Li	$\text{Li}^+ + \text{e}^- \rightarrow \text{Li(s)}$	- 3.05 V
K	$\text{K}^+ + \text{e}^- \rightarrow \text{K(s)}$	- 2.93 V
Ba	$\text{Ba}^{+2} + 2\text{e}^- \rightarrow \text{Ba}$	- 2.91 V
Ca	$\text{Ca}^{+2} + 2\text{e}^- \rightarrow \text{Ca(s)}$	- 2.87 V
Na	$\text{Na}^+ + \text{e}^- \rightarrow \text{Na(s)}$	- 2.71 V
Mg	$\text{Mg}^{+2} + 2\text{e}^- \rightarrow \text{Mg(s)}$	- 2.36 V
Al	$\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$	- 1.66 V
Mn	$\text{Mn}^{+2} + 2\text{e}^- \rightarrow \text{Mn}$	- 1.18 V
H <sub>2</sub> O	$\text{H}_2\text{O(l)} + \text{e}^- \rightarrow 1/2 \text{H}_2 + \text{OH}^-$	- 0.828 V
Zn	$\text{Zn}^{+2} + 2\text{e}^- \rightarrow \text{Zn(s)}$	- 0.76 V
Cr	$\text{Cr}^{+3} + 3\text{e}^- \rightarrow \text{Cr(s)}$	- 0.74 V
Fe	$\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$	- 0.44 V
Cd	$\text{Cd}^{+2} + 2\text{e}^- \rightarrow \text{Cd(s)}$	- 0.40 V
Co	$\text{Co}^{2+} + 2\text{e}^- \rightarrow \text{Co}$	- 0.28 V
Ni	$\text{Ni}^{+2} + 2\text{e}^- \rightarrow \text{Ni(s)}$	- 0.25 V
Sn	$\text{Sn}^{+2} + 2\text{e}^- \rightarrow \text{Sn(s)}$	- 0.14 V
Pb	$\text{Pb}^{+2} + 2\text{e}^- \rightarrow \text{Pb(s)}$	- 0.13 V
H <sub>2</sub>	$2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2(\text{g})$	0.00 V
Cu	$\text{Cu}^{2+} + 2\text{e}^- \rightarrow \text{Cu(s)}$	0.34 V
I <sub>2</sub>	$\text{I}_2 + 2\text{e}^- \rightarrow 2\text{I}^-$	0.54 V
Fe	$\text{Fe}^{3+} + \text{e}^- \rightarrow \text{Fe}^{2+}$	0.77 V
Hg	$\text{Hg}_2^{2+} + 2\text{e}^- \rightarrow \text{Hg(l)}$	0.79 V
Ag	$\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}$	0.80 V
Hg	$\text{Hg}^{2+} + 2\text{e}^- \rightarrow \text{Hg(l)}$	0.85 V
Br <sub>2</sub>	$\text{Br}_2 + 2\text{e}^- \rightarrow 2\text{Br}^-$	1.09 V
Pt	$\text{Pt}^{+2} + 2\text{e}^- \rightarrow \text{Pt}$	1.20 V
O <sub>2</sub>	$1/2 \text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow \text{H}_2\text{O}(\ell)$	1.23 V
Cl <sub>2</sub>	$\text{Cl}_2 + 2\text{e}^- \rightarrow 2 \text{Cl}^-$	1.36 V
Au	$\text{Au}^{+3} + 3\text{e}^- \rightarrow \text{Au(s)}$	1.40 V
F <sub>2</sub>	$\text{F}_2 + 2\text{e}^- \rightarrow 2\text{F}^-$	2.87 V



Increasing strength of reducing agent



Increasing strength of oxidising agent

# Electrochemical Constants

**D Siedentop**



## **Electrochemical Constants:**

## **Electrochemical Constants** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Electrochemical Constants**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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