

# Electricity: Fundamentals and Applications

This presentation delves into the fascinating world of electricity, covering its fundamental concepts and practical applications. We'll explore the nature of electric charge, the principles of electric fields and potential, and the behavior of electric currents and circuits. We will also discuss the important concepts of electromagnetic induction and transformers, which are essential for understanding how electricity is generated and used in modern technology.

 by Ayush Agrawal



## Electric Charge and Electric Field

### Electric Charge

Electric charge is a fundamental property of matter, responsible for electric and magnetic phenomena. There are two types of charges: positive and negative. Like charges repel each other, while unlike charges attract.

### Electric Field

An electric field surrounds any charged object, exerting a force on other charges within its vicinity. It is represented by lines of force, which point in the direction of the force on a positive charge.

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## Coulomb's Law and Electric Field Intensity

# Electricity Fundamentals And Applications

**R Sandford**



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web belly dance such as the bolro bandari chobiyyah and debke drummers use the drum note melody as an iterative foundation upon which to elaborate ornament and to interact with the dancer dancers use the rhythm to guide the form of their movements the elaborations and ornamentations are a dialog between drummer and dancer the link

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$   
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $137.1 \text{ g } \text{H}_2\text{O} / 149.4 \text{ g } \text{H}_2\text{O} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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 4 Al + 3  $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$   
 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times 1 \text{ mol Al} / 27 \text{ g Al} = 5.926 \text{ mol Al}$   
 $5.926 \text{ mol Al} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}} = 2.963 \text{ mol } \text{Al}_2\text{O}_3$   
 $2.963 \text{ mol } \text{Al}_2\text{O}_3 \times 101.964 \text{ g } \text{Al}_2\text{O}_3 / \text{mol } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$   
 c determine the percent yield  
 $\text{yield} = \frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighed by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h<sub>2</sub> are left over 13

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