



# El Viejo Carpintero

**R Bogdan**



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web write the oxidation and reduction reactions for each redox reaction the first one is done for you  
 $23 \text{ Fe} + \text{CO} \rightarrow \text{CO}_2 + \text{Fe}$   
 oxidation  $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$  reduction  $\text{Fe} + 2\text{e}^- \rightarrow \text{Fe}$   
 $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}_3 + 3 \text{ Ag}$  oxidation reduction  
 $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb}_2 + \text{Cu}$  oxidation reduction  
 $26 \text{ O}_2$

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 $\text{I}_2 + \text{O}_5 \rightarrow \text{I}_2\text{O}_5$   
 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$   
 $5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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 $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr} \rightarrow \text{Sr}^2+$  oxidized reducing agent  $\text{O}_2 \rightarrow \text{O}^{2-}$  reduced  
 $\text{Ox Ag} + 2\text{Li} \rightarrow \text{Li}_2\text{S} + \text{Li}$  oxidized red ag  
 $\text{S}_0 \rightarrow \text{S}_2$  reduced  
 $\text{Ox Ag} + 3\text{Cs} \rightarrow \text{Br}_2$

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