

## Answer Sheet for Sample Problems for Chemistry Level 1 Final Exam 2013-2014 Study Guide

### Electrons in Atoms – Chapter 5

1. What is the frequency of green light, which has a wavelength of  $4.90 \times 10^{-7} \text{ m}$ ?

$$c = \lambda \nu \quad \lambda = 4.90 \times 10^{-7} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{4.90 \times 10^{-7} \text{ m}}$$

$$\underline{6.12 \times 10^{14} \text{ Hz}}$$

2. An x-ray has a wavelength of  $1.15 \times 10^{-10} \text{ m}$ . What is its frequency?

$$c = \lambda \nu \quad \lambda = 1.15 \times 10^{-10} \text{ m}, c = 3.00 \times 10^8 \text{ m/s} \quad \nu = ? \text{ Hz} \quad \nu = \frac{c}{\lambda} = \frac{3.00 \times 10^8 \text{ m/s}}{1.15 \times 10^{-10} \text{ m}}$$

$$\underline{2.61 \times 10^{18} \text{ Hz}}$$

3. Which type of electromagnetic radiation has the most energy: Microwaves, X-Rays, or Visible light? Why?

**Ans: X-Rays, b/c they have the shortest wavelength and highest frequency of the types listed and frequency and energy are directly related.**

4. What is the energy of a  $9.50 \times 10^{13} \text{ Hz}$  radiation?

$$E = h\nu \quad \nu = 9.50 \times 10^{13} \text{ Hz}, h = 6.626 \times 10^{-34} \text{ J s} \quad E = ? \text{ J}$$

$$E = h\nu = (6.626 \times 10^{-34} \text{ J s}) (9.50 \times 10^{13} \text{ s}^{-1}) = \underline{6.29 \times 10^{-19} \text{ J}}$$

5. The element with the electron configuration of  $[\text{Kr}]5s^24d^{10}5p^1$  is indium (In)

6. Barium (Ba)

7. Write out the electron configurations for the following elements: (For the exam make sure you know how to do both the complete and the noble gas short cut for electron configurations)

a. Tc                      **Ans:  $1s^22s^22p^63s^23p^44s^23d^{10}4p^65s^24d^5$  or  $[\text{Kr}]5s^24d^5$**

b. Aluminum            **Ans:  $1s^22s^22p^63s^23p^1$  or  $[\text{Ne}] 3s^23p^1$**

8. Why are Cr and Cu exceptions to writing the electron configurations of elements?

**They are exceptions because instead of following the pattern of filling up the d orbitals Cr actually has a half filled s and d orbitals and Cu has a half filled s and completely filled d orbitals. The reason why this occurs is because the electrons in the d-sublevel are more stable with this electron configuration than the expected configuration.**

9. a. Ca

b. As

c. Sn

a. 2    Ca.

b. 5    .As.

c. 4    . Sn .

### Periodic Table and Periodic Trends – Chapter 6

1. Compare & contrast how Mendeleev's periodic table was organized to how Moseley organized it.

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**L Reisser**



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