

Chemistry: 2nd Semester Exam

Select the correct oxidation numbers for the following elements.

- Oxygen in O_2
a. -2 b. +2 c. -1 d. +1 e. 0
- Sodium in $NaCl$
a. -2 b. +2 c. -1 d. +1 e. 0
- Phosphorus in H_3PO_4
a. -5 b. -3 c. 0 d. +3 e. +5
- H in H_2O
a. -2 b. +2 c. -1 d. +1 e. 0
- O in H_2O
a. -2 b. +2 c. -1 d. +1 e. 0
- In which substance does bromine have an oxidation number of +1?
a. Br_2 b. HBr c. $HBrO$ d. $HBrO_2$
- Which species is the oxidizing agent in the following reaction? $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$
a. Cl_2 b. $2I^-$ c. I_2 d. $2Cl^-$
- Consider the following: $2MnO_4^- + 5H_2SO_3 \rightarrow 2Mn^{2+} + 3H_2O + 5SO_4^{2-} + 4H^+$
Which of the following elements does not experience a change in oxidation number?
a. S in H_2SO_3
b. O in MnO_4^-
c. Mn in MnO_4^-
- In which of the following unbalanced equations does chromium undergo oxidation?
a. $Cr^{3+} \rightarrow Cr$
b. $Cr^{2+} \rightarrow Cr^{3+}$
c. $Cr^{3+} \rightarrow Cr_2O_7^{2-}$
d. $CrO_4^{2-} \rightarrow Cr_2O_7^{2-}$
- In a redox reaction, the species which loses electrons is _____.
a. reduced b. oxidized c. neutralized
- Consider the following: $Cl_2 \rightarrow ClO_2^-$
The balanced half reaction is:
a. $Cl_2 + 3H_2O \rightarrow ClO_2^- + 6H^+ + 5e^-$
b. $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 4e^-$
c. $Cl_2 + 6H_2O + 2e^- \rightarrow 2ClO_2^- + 6H_2$
d. $Cl_2 + 6H_2O \rightarrow 2ClO_2^- + 12H^+ + 10e^-$

Label the Acid, Base, Conjugate Acid, and Conjugate Base



- $HC_2H_3O_2$ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- H_2O serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- H_3O^+ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base
- $C_2H_3O_2^-$ serves as the
a. Acid b. Base c. Conjugate Acid d. Conjugate Base

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M Walker



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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells

filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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