

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH from the concentration of either ion.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (page 714) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (page 714) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (page 714) | 4. $\text{pH} + \text{pOH} = 14.00$ (page 714) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
2. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/weak base)
3. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
4. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/weak base)
5. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
6. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
7. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
8. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
9. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)
10. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (strong acid/strong base)

11. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

12. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

13. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

14. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

15. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

16. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

17. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

18. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

19. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

20. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

21. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

19 2 Section Review Hydrogen Ions And Acidity

Roman Wölfel



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Table of Contents 19 2 Section Review Hydrogen Ions And Acidity

1. Understanding the eBook 19 2 Section Review Hydrogen Ions And Acidity
 - The Rise of Digital Reading 19 2 Section Review Hydrogen Ions And Acidity
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Section Review Hydrogen Ions And Acidity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Section Review Hydrogen Ions And Acidity
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Section Review Hydrogen Ions And Acidity
 - Personalized Recommendations
 - 19 2 Section Review Hydrogen Ions And Acidity User Reviews and Ratings
 - 19 2 Section Review Hydrogen Ions And Acidity and Bestseller Lists
5. Accessing 19 2 Section Review Hydrogen Ions And Acidity Free and Paid eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity Public Domain eBooks
 - 19 2 Section Review Hydrogen Ions And Acidity eBook Subscription Services
 - 19 2 Section Review Hydrogen Ions And Acidity Budget-Friendly Options
6. Navigating 19 2 Section Review Hydrogen Ions And Acidity eBook Formats

- ePub, PDF, MOBI, and More
 - 19 2 Section Review Hydrogen Ions And Acidity Compatibility with Devices
 - 19 2 Section Review Hydrogen Ions And Acidity Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Section Review Hydrogen Ions And Acidity
 - Highlighting and Note-Taking 19 2 Section Review Hydrogen Ions And Acidity
 - Interactive Elements 19 2 Section Review Hydrogen Ions And Acidity
 8. Staying Engaged with 19 2 Section Review Hydrogen Ions And Acidity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Section Review Hydrogen Ions And Acidity
 9. Balancing eBooks and Physical Books 19 2 Section Review Hydrogen Ions And Acidity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Section Review Hydrogen Ions And Acidity
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 19 2 Section Review Hydrogen Ions And Acidity
 - Setting Reading Goals 19 2 Section Review Hydrogen Ions And Acidity
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 19 2 Section Review Hydrogen Ions And Acidity
 - Fact-Checking eBook Content of 19 2 Section Review Hydrogen Ions And Acidity
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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