

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. acid dissociation constant (K_a) | 6. hydronium ion concentration |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ (Eq. 14.1) | 2. $K_b = \frac{[\text{B}][\text{H}_2\text{O}]}{[\text{BH}^+]}$ (Eq. 14.2) |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems that require calculations.

1. What products are **1** **2** formed by the autoionization of water? Write the chemical formula of the products of these autoionizations and the chemical formula of the water molecules that are produced in this process. **3** **4**
2. Hydrochloric acid dissociates into **1** **2** in water. Write the **3** **4** dissociation equation for this acid. Identify **5** **6** **7** **8** **9** **10** the conjugate acid-base pairs.
3. **1** **2** **3** **4** **5** **6** **7** **8** **9** **10** **11** **12** **13** **14** **15** **16** **17** **18** **19** **20** **21** **22** **23** **24** **25** **26** **27** **28** **29** **30** **31** **32** **33** **34** **35** **36** **37** **38** **39** **40** **41** **42** **43** **44** **45** **46** **47** **48** **49** **50** **51** **52** **53** **54** **55** **56** **57** **58** **59** **60** **61** **62** **63** **64** **65** **66** **67** **68** **69** **70** **71** **72** **73** **74** **75** **76** **77** **78** **79** **80** **81** **82** **83** **84** **85** **86** **87** **88** **89** **90** **91** **92** **93** **94** **95** **96** **97** **98** **99** **100**

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2. H_3O^+ and OH^-
3. H_2O
4. H_3O^+ and OH^-
5. H_2O
6. H_3O^+
7. OH^-
8. H_2O
9. H_3O^+
10. OH^-
11. H_2O
12. H_3O^+
13. OH^-
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**Thirumalaisamy P.
Velavancorresponding**



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