

# X 220

## Excavator

### Operation & Maintenance Manual

(S/N 508215001 & Above)



# X220 Service Manual

**Huangqi Zhang**



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the concentrations of all species at equilibrium check your final answers with a  $q$  calculation let  $x$  be the number of mol l of  $h_2$  or  $i_2$  that is lost to reach equilibrium  $h_2(g) + i_2(g) \rightleftharpoons 2hi(g)$

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web as a result of the reaction the equilibrium concentration of  $nh_3(g)$  became  $0.540\text{ M}$  what is the value of the equilibrium constant for this reaction at the given temperature of  $700\text{ C}$   $n_2(g) + 3h_2(g) \rightleftharpoons 2nh_3(g)$  i write the initial concentrations of reactants and products

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web equilibrium  $1.864\text{ M}$   $1.364\text{ M}$   $0.636\text{ M}$  equilibrium  $1.600\text{ M}$   $1.000\text{ M}$   $0.400\text{ M}$  trial 5 a  $g$  b  $g$  c  $g$  trial 6 a  $g$  b  $g$  c  $g$  initial  $1.400\text{ M}$   $1.200\text{ M}$   $0.400\text{ M}$  initial  $0.750\text{ M}$   $2.000\text{ M}$   $0.250\text{ M}$  change change equilibrium  $1.388\text{ M}$   $1.188\text{ M}$   $0.412\text{ M}$  equilibrium  $0.675\text{ M}$   $1.925\text{ M}$   $0.325\text{ M}$  1 examine model 1 a

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