

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentration, molar concentration, and pH of weak and strong acids and bases.
3. Calculate the pH of weak (or strong) acids.

## Vocabulary

- |                   |                                     |                     |
|-------------------|-------------------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair         | 3. proton indicator |
| 4. autoionization | 5. hydroxide ion (OH <sup>-</sup> ) | 6. hydronium ion    |
|                   | 7. weak indicator                   | 8. pH               |

## Key Equations

- |  |  |
|--|--|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ (Eq. 14.1) | 2. $K_b = \frac{[\text{B}][\text{H}_2\text{O}]}{[\text{BH}^+]}$ (Eq. 14.2) |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.3)                      | 4. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.4)                            |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid/base.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (autoionization of water)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (autoionization of water)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (autoionization of water)
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7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (autoionization of water)
8.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (autoionization of water)
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9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
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