

# Chemistry: 2<sup>nd</sup> Semester Exam

Select the correct oxidation numbers for the following elements.

- Oxygen in  $O_2$   
a. -2                      b. +2                      c. -1                      d. +1                      e. 0
- Sodium in  $NaCl$   
a. -2                      b. +2                      c. -1                      d. +1                      e. 0
- Phosphorus in  $H_3PO_4$   
a. -5                      b. -3                      c. 0                      d. +3                      e. +5
- H in  $H_2O$   
a. -2                      b. +2                      c. -1                      d. +1                      e. 0
- O in  $H_2O$   
a. -2                      b. +2                      c. -1                      d. +1                      e. 0
- In which substance does bromine have an oxidation number of +1?  
a.  $Br_2$                       b.  $HBr$                       c.  $HBrO$                       d.  $HBrO_2$
- Which species is the oxidizing agent in the following reaction?  $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$   
a.  $Cl_2$                       b.  $2I^-$                       c.  $I_2$                       d.  $2Cl^-$
- Consider the following:  $2MnO_4^- + 5H_2SO_3 \rightarrow 2Mn^{2+} + 3H_2O + 5SO_4^{2-} + 4H^+$   
Which of the following elements does not experience a change in oxidation number?  
a. S in  $H_2SO_3$   
b. O in  $MnO_4^-$   
c. Mn in  $MnO_4^-$
- In which of the following unbalanced equations does chromium undergo oxidation?  
a.  $Cr^{3+} \rightarrow Cr$   
b.  $Cr^{2+} \rightarrow Cr^{3+}$   
c.  $Cr^{3+} \rightarrow Cr_2O_7^{2-}$   
d.  $CrO_4^{2-} \rightarrow Cr_2O_7^{2-}$
- In a redox reaction, the species which loses electrons is \_\_\_\_\_.  
a. reduced                      b. oxidized                      c. neutralized
- Consider the following:  $Cl_2 \rightarrow ClO_2^-$   
The balanced half reaction is:  
a.  $Cl_2 + 3H_2O \rightarrow ClO_2^- + 6H^+ + 5e^-$   
b.  $Cl_2 + 3H_2O \rightarrow 2ClO_2^- + 6H^+ + 4e^-$   
c.  $Cl_2 + 6H_2O + 2e^- \rightarrow 2ClO_2^- + 6H_2$   
d.  $Cl_2 + 6H_2O \rightarrow 2ClO_2^- + 12H^+ + 10e^-$

## Label the Acid, Base, Conjugate Acid, and Conjugate Base



- $HC_2H_3O_2$  serves as the  
a. Acid                      b. Base                      c. Conjugate Acid                      d. Conjugate Base
- $H_2O$  serves as the  
a. Acid                      b. Base                      c. Conjugate Acid                      d. Conjugate Base
- $H_3O^+$  serves as the  
a. Acid                      b. Base                      c. Conjugate Acid                      d. Conjugate Base
- $C_2H_3O_2^-$  serves as the  
a. Acid                      b. Base                      c. Conjugate Acid                      d. Conjugate Base

# 2nd Semester Exam Chemistry Form

**R Sandford**



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