

# THE "BIG FIVE" PERSONALITY FACTORS

(Memory tip: Picturing a CANOE will help you recall these.)

Trait Dimension	Endpoints of the Dimension		
<b>Conscientiousness</b>	Organized	↔	Disorganized
	Careful	↔	Careless
	Disciplined	↔	Impulsive
<b>Agreeableness</b>	Soft-hearted	↔	Ruthless
	Trusting	↔	Suspicious
	Helpful	↔	Uncooperative
<b>Neuroticism</b> (emotional stability vs. instability)	Calm	↔	Anxious
	Secure	↔	Insecure
	Self-satisfied	↔	Self-pitying
<b>Openness</b>	Imaginative	↔	Practical
	Preference for variety	↔	Preference for routine
	Independent	↔	Conforming
<b>Extraversion</b>	Sociable	↔	Retiring
	Fun-loving	↔	Sober
	Affectionate	↔	Reserved

# Quizlet Ap Psych Chapter 4

**RM Cervero**



## **Quizlet Ap Psych Chapter 4:**

## Reviewing **Quizlet Ap Psych Chapter 4**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Quizlet Ap Psych Chapter 4**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}^{2+} + \text{O}_2 \rightarrow \text{Cr}^{3+} + \text{H}_2\text{O}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^-$   $2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{AuCN} + 2\text{e}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} + 3\text{O}_2$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh ions or the oh h<sub>2</sub>o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo and cr oh<sub>4</sub> in basic solution created by jay

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web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$  the above equation takes place in two

stages first bromine reacts with  $\text{OH}^-$  ions to form bromide ions and  $\text{BrO}^-$  ions second the  $\text{BrO}^-$  ions react to form bromide ions and bromate ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{H}^+ \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+} + \text{H}_2\text{O}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te}$  10  $\text{ReO}_4^- + \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  
c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and