

18.4 Acid-Base Chemistry: Theory

Multiple-Choice Questions

1. Which combination of quantities are dimensional quantities? (correct)
- (A) pressure, concentration, pH, amount of substance
 (B) volume, molarity, pH, amount of substance
 (C) volume, molarity, pH, concentration of ions
 (D) amount of substance, pH, concentration of ions, concentration of substance

(A) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (B) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

(C) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (D) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$ (E) $\text{pH} = \frac{[\text{H}^+]}{10^{-7}}$ (F) $\text{pH} = \frac{[\text{H}^+]}{10^{-14}}$

2. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $\text{pH} = 4.5$ (B) $\text{pH} = 5.5$ (C) $\text{pH} = 6.5$ (D) $\text{pH} = 7.5$ (E) $\text{pH} = 8.5$

(F) $\text{pH} = 9.5$ (G) $\text{pH} = 10.5$ (H) $\text{pH} = 11.5$ (I) $\text{pH} = 12.5$

3. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)

(A) $\text{pH} = 4.5$ (B) $\text{pH} = 5.5$ (C) $\text{pH} = 6.5$ (D) $\text{pH} = 7.5$

4. Which of the following is a weak acid? (correct)



5. A solution of an acid has a pH of 2.5. The concentration of the acid is 0.1 mol/L. What is the pK_a of the acid? (correct)

(A) $\text{pK}_a = 2.5$ (B) $\text{pK}_a = 3.5$ (C) $\text{pK}_a = 4.5$ (D) $\text{pK}_a = 5.5$ (E) $\text{pK}_a = 6.5$

(F) $\text{pK}_a = 7.5$ (G) $\text{pK}_a = 8.5$ (H) $\text{pK}_a = 9.5$ (I) $\text{pK}_a = 10.5$

6. An acid-base indicator has a pK_a of 4.5. In a solution of the indicator, the ratio of the concentration of the acid form to the concentration of the base form is 10:1. What is the pH of the solution? (correct)



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